CAMBRIDGE INTERNATIONAL EXAMINATIONS

International General Certificate of Secondary Education

MARK SCHEME for the May/June 2014 series

0654 CO-ORDINATED SCIENCES

0654/62 Paper 6 (Alternative to Practical), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the May/June 2014 series for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level components and some Ordinary Level components.



		. J -		IGCSE – May/June 2014	0654	62
<u> </u>				1903E — May/Julie 2014	U03 4	UZ
1	(a)	(pu	rple d	lue to) pH above 8/alkaline (conditions);		[1]
	(b)	128 728	ading for time taken with units (allow units in table); Bis for block A ; Is for block B ; It is for block B ;			[3]
	(c)	(i)	diffu	sion ; (NOT osmosis)		[1]
		(ii)	redu	ices pH/takes pH below 8 (so it goes colourless);		[1]
	(d)	(i)	(B q	uicker as) smaller distance/volume/size/surface a	rea/other correct	; [1]
		(ii)		oli (walls)/lungs/capillaries one cell thick/large ter path(way);	surface (area)/th	nin / [1]
	(e)	(i)	diffe	rent sized blocks/greater range of block sizes;		[1]
		(ii)	<u>time</u>	ither axis: and volume/(surface) area/dimensions/size; (idrawn)	gnore units, and	any [1] [Total: 10]
2	(a)	(i)	carl	oonate / CO ₃ ²⁻ ;		[1]
		(ii)	(aqu	er order: leous) silver nitrate/AgNO ₃ /lead nitrate/Pb(NO ₃) ₂ ; c acid/HNO ₃ ;		[2]
		(iii)	exot	hermic ;		[1]
	(b)	(i)	copp iron(per/Cu ²⁺ ; (II)/Fe ²⁺ ;		[2]
		(ii)		tion diagram must <u>see</u> both funnel and paper ; relevant labels ;		[2]
		(iii)		ens/(turns) brown ; ation ;		[2]

Mark Scheme

Syllabus

Paper

[Total: 10]

Page 2

Page 3		Mark Scheme	Syllabus	Paper
		IGCSE – May/June 2014	0654	62
(a) (i)	0.14 1.3 <u>0</u>	• •		[2]
` ,	0.38 0.29 0.23			
	0.18 0.15	(ecf) ;;		[2]
(iii)	(lam	p is) less bright/dimmer ;		[1]
. , . ,	0.09 0.04 0.02 0.01	(0.181) (0.086) (0.038) (0.022) (0.015) ;; correct = 2 marks, one error = 1 mark BUT max 1 if	any rounding error)	[2]
(ii)		ght line, positive slope ; <u>ing</u> through origin ;		[2]
(iii)	disaç	gree (no mark)		
	$\frac{V}{l}$ n	ot constant/as length/ l increases, V decreases ;		[1] [Total: 10]
				[101411 10]
	cess - spirat lanati	<u>tion</u> ;		[1]
		<u>ion</u> of water (at mesophyll cells) ; ater vapour from leaves (through stomata)/water <u>q</u> i	iven off by leaves ;	[1]
reco timin repe othe	ord stang/useats/l	f varying wind speed e.g. hairdryer/fan ; art/end distance ; se of a stopclock; more than one experiment ; one) conditions constant e.g. same plant, plant size imental method not the effect) ;	e, temp, light (looking	[max 4]
(c) (i)		ing from left, right or middle of bubble (1.0, 1.5 or 3.0 or 3.5) at end;	2.0 at start) to match	h [1]
(ii)		high) ; low) ; (ecf)		[2]
	/ironn nidity	mental) temp ; ;		
wate rain		ailability ;		[max 1]
				[Total: 10]

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Page 4	Mark Scheme	Syllabus	Paper
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5 note: for part (a) and part (b)(i) allow letter or name

(a) add A to B will produce/gas/bubbles/CO₂ therefore C is BaCl₂;
 add B to C will produce(white) ppt therefore B is Na₂CO₃;
 therefore A must be HCl;
 (or any other way)

(b) (i) A and B (either order) or names; [1]

(ii) evaporation; [1]

(iii) diagram; (allow a 'series' of diagrams to show evaporation in a beaker) two relevant labels; [2]

(c) use of sodium hydroxide (aq) and/or (aq) ammonia; white ppt; dissolves in excess/(solution) turns colourless; [3] (if WRONG reagent, maximum mark 1 for white ppt)

[Total: 10]

6 (a) (i) 4.5;

(ii) 3600; [1]

(iii) $4.5 \times 12 \times 3600$ (ecf); 194400; [2]

(b) (i) 83 °C; 63 °C (ecf);

(ii) $0.5 \times 4200 \times 63 \text{ (ecf)}$; [2] 132300 (J);

(c) efficiency = $\frac{\text{useful (energy) out}}{\text{total (energy) in}}$ (× 100 %);

$$\frac{132\,300}{194\,400} = 68\% \text{ (ecf)};$$

[Total: 10]